Tina Yetukuri

Per:7

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**Gas Law Photo Quiz**

A febreze spray can is a real life application of Boyle’s Law. Boyle’s law states that pressure and volume are inversely related. Aerosol cans have very high pressure in them due to the liquefied gas. This pressure is released or decreased when the nozzle is pressed and the product comes out of the can into the atmosphere, where there is an increase in volume. Boyle’s law takes place at a constant temperature and fixed mass. The variables that are changing are pressure and volume. As the volume decreases, the pressure increases because there is less space for the particles and hitting the edge of the container more often at high speed. And inversely as volume increases, pressure decreases because there is more room for the particles to freely move around and hit the surrounding container less often. The formula for Boyle’s law is

***P1V1 = P2V2.***

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